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General Questions

Practice Questions

Q1. What is the electron configuration of a neutral oxygen atom?

A. 1s2 2s2 2p4
B. 1s2 2s2 2p6
C. 1s2 2s2 2p3
D. 1s2 2s2 2p5

Solution: Oxygen has 8 electrons, and its electron configuration is 1s2 2s2 2p4.

Q2. What is the principal quantum number for the outermost electrons in a potassium atom?

A. 1
B. 2
C. 3
D. 4

Solution: Potassium (K) has an atomic number of 19, and its outermost electrons are in the n=4 level.

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