



SoulShift - Educational Q&A Platform

General Questions

Practice Questions



Q1. What is the electron configuration of a neutral oxygen atom?

- A. $1s^2 2s^2 2p^4$
- B. $1s^2 2s^2 2p^6$
- C. $1s^2 2s^2 2p^3$
- D. $1s^2 2s^2 2p^5$

Solution: Oxygen has 8 electrons, and its electron configuration is $1s^2 2s^2 2p^4$.

Q2. What is the principal quantum number for the outermost electrons in a potassium atom?

- A. 1
- B. 2
- C. 3
- D. 4

Solution: Potassium (K) has an atomic number of 19, and its outermost electrons are in the $n=4$ level.



